Additions and Corrections

Solvent Mass Transfer across Ion-Exchange Membranes [*J. Am. Chem. Soc.*, **88**, 661 (1966)]. By A. S. TOMBALAKIAN, M. WORSLEY, and W. F. GRAYDON, Department of Chemical Engineering and Applied Chemistry, University of Toronto, Toronto, Ontario, Canada.

The interaction coefficients reported in Table V were incorrect. The corrected table is printed below.

Table V. Ion-Water and Water-Membrane Interaction Coefficients

Membrane no.	Ion-pair exchange system	Interaction coefficients, dynes sec cm^{-1} × 10^{13} f_{IW}*	f_{WW}
1	K^+–H^+	K^+, 0.89	1.7
	Na^+–H^+	Na^+, 1.5	1.5
	Li^+–H^+	Li^+, 2.3	0.84
	$(CH_3)_4N^+$–H^+	$(CH_3)_4N^+$, 2.8	0.41
	$(C_2H_5)_4N^+$–H^+	$(C_2H_5)_4N^+$, >3.4	<0

Oxidative Dimerization of Carbanions in Liquid Ammonia [*J. Am. Chem. Soc.*, **89**, 3659 (1967)]. By EDWIN M. KAISER, Department of Chemistry, University of Missouri, Columbia, Missouri 65201.

The correct formula for compound **2** in eq 1 should be $(C_6H_5)_2CHCH(C_6H_5)_2$.

Photochemical Formation of a Substituted Bicyclo[1.1.1]pentane [*J. Am. Chem. Soc.*, **89**, 6376 (1967)]. By ALBERT PADWA and EDWARD ALEXANDER, Department of Chemistry, State University of New York at Buffalo, Buffalo, New York 14214.

On page 6377, column 1, line 28, $\Phi_{III} = 0.004$ should read $\Phi_{II} = 0.006$.

Equilibration Studies. The Energy Differences for Some Six-Membered Heterocyclic Methyl Amide–Imidate Isomer Pairs [*J. Am. Chem. Soc.*, **90**, 1569 (1968)]. By PETER BEAK, JAMES BONHAM, and JAMES T. LEE, JR., W. A. Noyes Chemical Laboratory, University of Illinois, Urbana, Illinois 61801.

On page 1570 in Table I, the column $\Delta G_g°$ should be deleted.

On page 1571 in Figure 1, the expressions involving G and $\Delta G°$ and in column 2, line 10, the sentence beginning "No free energy. . ." should be deleted.

On page 1572, column 1, line 1, delete the sentence beginning "The limiting values. . ."

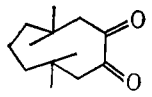
In Table VII, the last energy under $\Delta H_l°$ should be −17.0.

Thallium in Organic Synthesis. III. Coupling of Aryl and Alkyl Grignard Reagents [*J. Am. Chem. Soc.*, **90**, 2423 (1968)]. By ALEXANDER MCKILLOP, L. F. ELSON, and EDWARD C. TAYLOR, Department of Chemistry, Princeton University, Princeton, New Jersey 08540.

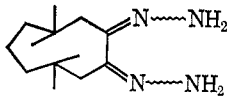
Reference 6 should read: H. Gilman and R. G. Jones, *J. Am. Chem. Soc.*, **62**, 2357 (1940).

A Rigid, Internally Saturated Derivative of Cyclononane [*J. Am. Chem. Soc.*, **90**, 3233 (1968)]. By ALFRED T. BLOMQUIST and ROBERT D. MILLER, Department of Chemistry, Cornell University, Ithaca, New York 14850.

The structures of compounds **1** and **2** shown on page 3233 are incorrect. They should be as shown below.



1



2

The Structure of the *trans*-Dioxotetracyanomolybdate(IV) Ion in the Crystalline Salt $NaK_3MoO_2(CN)_4 \cdot 6H_2O$ [*J. Am. Chem. Soc.*, **90**, 3374 (1968)]. By V. W. DAY and J. L. HOARD, Department of Chemistry, Cornell University, Ithaca, New York 14850.

On page 3374, in line 6 of the abstract, and in the headings of both Table I and Table II, the chemical formula of the salt should be $NaK_3MoO_2(CN)_4 \cdot 6H_2O$.

On page 3377, last paragraph, the first sentence should read, "The MoCN chain length (or Mo–N distance) of 3.372 Å. . . ."

Book Reviews

The Chemistry of Organometallic Compounds. The Main Group Elements. By JOHN J. EISCH, the Catholic University of America. The Macmillan Co., 866 Third Ave., New York, N. Y. 1967. xi + 178 pp. 16 × 24 cm. $5.95.

This brief book provides a unique treatment of its subject. As the author states in the introduction, the emphasis is on the organic behavior of the group R in different RM compounds, and on the reactivity of such organometallics toward organic substrates This is, then, very much an organic chemist's view of organometallic chemistry.

The introductory pages are followed by an excellent chapter on the behavior of metals toward organic systems, wherein all such